

LVII.—*The Determination of the Dissociation Pressures of Hydrated Salts by a Dynamical Method.*

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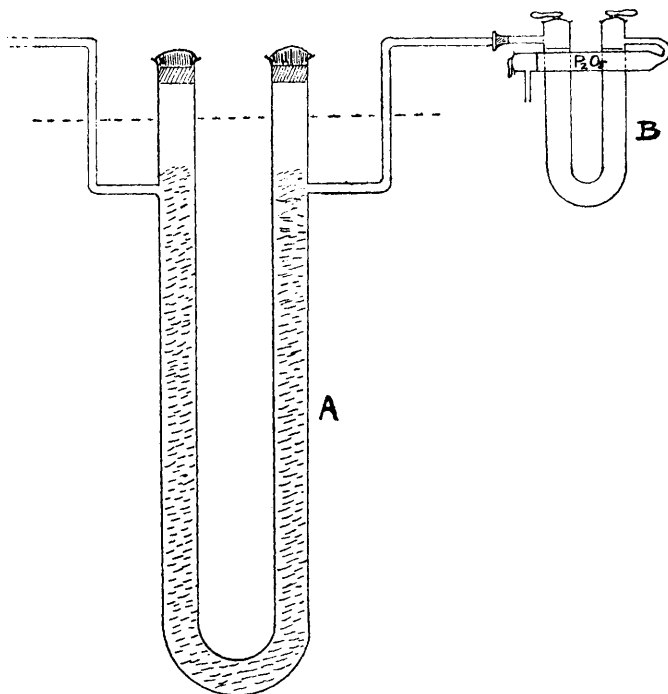
PROBABLY no branch of physico-chemical investigation has aroused a more lively and sustained interest than that dealing with the dehydration of the so-called molecular compounds of water with substances which are electrolytically dissociated in aqueous solution

He adopted Tammann's explanation: "Schwer zu vermeiden war ferner eine Störung, die wohl der Adsorption, dem Festhalten einer Lösungshaut an den grossen Oberflächen der gepulverten Salze, zuzuschreiben ist. Nach frischer Beschickung des Apparates traten zunächst viel zu grosse Drucke auf."

EXPERIMENTAL.

The author has carried out a number of experiments by the transpiration method, but instead of measuring the volume of air

FIG. 1.



aspirated through the apparatus, which involves troublesome corrections for changes of barometric pressure, the air, after passing over the hydrate, was allowed to bubble through water at the same temperature. If w_1 , w_2 are the weights of water removed from the hydrate and pure water respectively, at any temperature θ , it is easily shown that:

$$p_{\theta} = \frac{w_1}{w_2} \pi_{\theta} \dots \dots \dots (1).$$

dynamic pressure is due to saturated solution adsorbed on the solid salt.

In conclusion, I desire to thank Dr. Lapworth for the interest he has shown during the course of the work.

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