

following is a formula taken from a recent issue of the "*Photographische Chronik*":

Water .....	1000 c.c.
Metol .....	7 gm.
Hydroquinone .....	8 gm.
Sodium Sulphite, dry .....	100 gm.
Ammonia (sp. gr. 0.923) .....	32 c.c.
Potassium Bromide .....	6 gm.

For use, dilute with four volumes of water.

The following developer was proposed about the time that the autochrome came into use:

Water .....	1000 c.c.
Sodium Sulphite, dry .....	40 gm.
Metol .....	6.5 gm.
Hydroquinone .....	2 gm.
Hypo. ....	0.1 gm.
Ammonia (sp. gr. 0.880) .....	20 c.c.
Potassium Bromide .....	2.5 gm.

For use, dilute with an equal volume of water.

For some lines of work, especially photomicrography, the second development may be omitted. Given a good exposure and a good first development, the oxidizing agent is then applied until the silver deposit is completely removed. The plate is then washed and dried. The unchanged silver bromide will act as the opaque coating, upon which the correct color effect depends.

Striking effects can be obtained by the use of autochromes in photographing the colors produced by the transmission of polarized light through rock section.

**Melting Point of Ammonium Sulphate.**—James Kendall and Arthur W. Davidson of Columbia University (*Journ. Ind. and Eng. Chem.* 1921 xiii, 303–304) have found the melting point of *acid* ammonium sulphate to be  $146.9^{\circ} \pm 0.5^{\circ} \text{C.}$ , and the melting point of *neutral* ammonium sulphate to be  $513^{\circ} \pm 2^{\circ} \text{C.}$  The extreme discrepancies in the literature concerning these melting points were found to be due essentially to the instability of the neutral salt when heated in an open tube.

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