

## Kinetics and Mechanism of Oxidation of Co (II) Ternary Complexes Involving N-(2-Acetamido) Iminodiacete and Some Amino Acids Acid by Periodate

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**Abstract :** The kinetics of oxidation of the cobalt (II) complexes,  $[\text{CoII}(\text{ADA})(\text{Gly})(\text{H}_2\text{O})_2]^-$ , (ADA = N-(2-acetamido) iminodiacetic acid and Gly = Glycine) by periodate in aqueous acetate medium to cobalt (III) have been studied spectrophotometrically at 530 nm over the 30–50°C and a variety pH 4.57-5.25 range and  $I = 0.50 \text{ mol dm}^{-3}$  under pseudo first order condition by taking large excess of oxidant  $[\text{IO}_4^-]$  and it obeys the following rate law:  $\text{Rate} = [\text{CoII}(\text{ADA})(\text{Gly})(\text{H}_2\text{O})_2]^- [\text{H}_5\text{IO}_6] \{k_4K_6 + (k_5K_7K_5/[\text{H}^+])\}$ . Also, the kinetics of oxidation of the cobalt(II) complexes,  $[\text{CoII}(\text{ADA})(\text{Val})(\text{H}_2\text{O})_2]^-$  (ADA = N-(2-acetamido) iminodiacetic acid and Val = valine) by periodate in aqueous medium to cobalt (III) have been studied spectrophotometrically at 580 nm over the 30–50°C and a variety pH 4.3-5.12 range and  $I = 0.50 \text{ mol dm}^{-3}$  under pseudo first order condition by taking large excess of oxidant  $[\text{IO}_4^-]$  and it obeys the following rate law:  $\text{Rate} = [\text{CoII}(\text{ADA})(\text{Val})(\text{H}_2\text{O})_2]^- [\text{H}_5\text{IO}_6] \{k_4K_6 + (k_5K_7K_5/[\text{H}^+])\}$

**Keywords :** periodate, oxidation, cobalt (II), glycine, valine acid, n-(2-acetamido imino-diacetato)

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