AS TO PLATINUM

Editor of the Journal of Industrial and Engineering Chemistry:

Much has been written of late, and more said, regarding the use of platinum in jewelry, and it has been broadly intimated that the jewelers are not living up to their agreement of last April with the Government. It is worth while to repeat the terms of this pledge of the Jewelers' Vigilance Committee.

"We pledge ourselves to discontinue and strongly recommend to all manufacturing and retail jewelers of the United States that they in a truly patriotic spirit discourage the manufacture, sale and use of platinum in all bulky and heavy pieces of jewelry.

"During the period of the war, or until the present supplies of platinum shall be materially augmented, we pledge ourselves to discontinue and recommend that the jewelry trade discourage the use of all non-essential platinum findings or parts of jewelry, such as scarfpin stems, pin tongues, joints, catches, swivels, spring rings, ear backs, etc., where gold would satisfactorily serve.

"Be it further resolved that the jewelry trade encourage by all means in its power, the use of gold in combination with platinum, wherever proper artistic results may be obtained."

Having been in close touch with the platinum situation I desire to state from personal knowledge my belief that the jewelers have fully lived up to their pledges and in many cases gone beyond them in efforts to conserve platinum.

It was the manufacturing jewelers who entered into this agreement, and it is not surprising that some retailers have been making great efforts to work off their stock on hand, and have thereby opened themselves to criticism.

Whether a metal, so limited in supply and so invaluable in scientific industry, ought to be used at all in jewelry is a fair question, but it is not the question in point. A large and legitimate platinum jewelry industry has sprung up in recent years, and the question is whether the exigencies of the present platinum situation demand the immediate wrecking of this industry by having the Government commandeer all platinum; personally, I do not believe that at present they do.

I hold no brief for the jewelers, but I think this statement should be made in fairness to them; and I may add that it is my conviction, that, should the Government be placed in straits from lack of platinum for the manufacture of war material, the jewelry trade can be relied on to find a way of furnishing all that is needed.

JAS. LEWIS HOWE

Special Committee on Platinum,
Chemistry Committee of the National
Research Council

WASHINGTON AND LEE UNIVERSITY
LEXINGTON, VIRGINIA

PLATINUM RESOLUTIONS

At the recent Pittsburgh meeting of Section C (Chemistry) of the American Association for the Advancement of Science, the following resolution was unanimously passed:

WHEREAS (1) The Government of the United States has purchased from Russia and safely brought to the United States twenty-one thousand (21,000) ounces of platinum, the largest amount that has ever been shipped to this country;

(2) The separation of so great an amount of platinum will offer scientific investigators an opportunity to study the chemical combinations and mineralogical associations of the platinum group of minerals,

Therefore it is *Resolved* (1) That Section C, the Section on Chemistry of the American Association for the Advancement of Science, respectfully request that the War Industries Board, the Bureau of Standards, the Bureau of Mines, the American Chemical Society, and others who are interested in chemistry, be offered the opportunity by the United States Government to coöperate with it in the separation of the platinum group of minerals of the above-mentioned material, and that the residue, of which there may be thousands of ounces, be loaned to such scientific investigators who can undoubtedly obtain interesting scientific results, and

(2) If necessary, that one platinum works be commandeered, with proper compensation, for a certain length of time so that the work of separation can be carried on with the greatest care and observation, and

(3) That as much as possible of this platinum be loaned to those who have need of platinum for chemical investigation, the platinum to belong to the Government of the United States and subject to the call of the Government when needed for raw or industrial purposes. As the expense of making crucibles or other utensils is small in comparison to the value of the platinum, this would offer a most unique opportunity to the chemists of the country.

To present this matter to the proper authorities a committee was appointed consisting of Dr. W. A. Noyes, *Chairman*, and Dr. W. F. Hillebrand.

NEW YORK CITY
January 5, 1918

GEORGE F. KUNZ

FUEL FOR MANUFACTURE OF CHEMICALS

Editor of the Journal of Industrial and Engineering Chemistry:

In view of the shortage in fuel supply and the great demand for large supplies of cheap fuel for the manufacture of chemicals and other products necessary for the war and agricultural purposes, I am bringing to your attention our unusual supply of natural gas at Shreveport, which is available to large industrial consumers on a basis that makes it cheaper than water power or that derived from the use of coal.

There is now available at Shreveport about one billion cubic feet of gas daily, and on the basis of scientific estimate the content of the Shreveport field is two trillion cubic feet, and a very small amount of this gas is now being utilized. This estimate includes only the field as already defined and does not consider other fields which are being opened up near by in drilling for oil. A prominent geologist has recently stated that the Shreveport field has the largest supply of natural gas to be found in the United States.

In addition to our gas supply, we have an abundance of raw materials that are needed for war industries at this time. There is close at hand an abundance of iron, petroleum, lignite, limestone, sulfur, and salt, and generous supplies of asphalt, gypsum, kaolin, sand, gravel, clay, etc. It ought to be especially noted at this time that Louisiana has the largest deposits of sulfur and salt to be found in the United States.

SHREVEPORT, LOUISIANA
November 21, 1917

ELLIS SMITH

A STUDY OF THE ESTIMATION OF FAT IN CONDENSED MILK, ETC.—CORRECTION

In our article printed under the above title [THIS JOURNAL, 9 (1917), 1111] the following changes should be made:

Page 1113, 1st col., line 18—"0.4" should read "0.04;" Table V, 3rd col., No. 8—"2.5005" should read "2.0505."

C. H. BIESTERFELD AND O. L. EVENSON

COMPOSITION OF LOGANBERRY JUICE AND PULP—CORRECTION

In the article under the above title [THIS JOURNAL, 9 (1917), 1043] note the following rearrangement of 5th line, Table I:

	I	II	III
Per cent Ash { Alkalinity.....	0.4139	0.5785	0.4226
{ K ₂ CO ₃	0.4130	0.5075	0.288
should read:			
Per cent Ash.....	0.4139	0.5785	0.4226
Alkalinity (as K ₂ CO ₃).....	0.4130	0.5075	0.288

M. R. DAUGHTERS