aration of various metals by electrolysis. Bernthsen was formerly a student and colleague of Bunsen, and he could speak with some authority. The mayor, Dr. Wilckens, accepted the monument in the name of the university city. Representatives of the German Chemical Society and a score of similar organizations deposited upon the monument laurel wreaths and other floral tributes suitably inscribed, as a mark of esteem for the man who so ably served his day and generation. Besides the scientific organizations, many of the German universities and technical schools sent their representatives and floral offerings to do honor to the occasion which proved a notable day in the history of the city. The exercises were largely attended, and several members of Bunsen's family were interested spectators.

Bunsen was not only one of the foremost scientists of his day, but he was a good citizen and an amiable man, beloved by all classes. He received many honors in his lifetime. The name of the street on which he lived during the ten years that he was professor *emeritus* was changed to Bunsen Strasse—rather an unusual honor, but one richly deserved.

Bunsen came to Heidelberg as professor at the age of forty-one. During the thirteen years previous, he had served the universities of Breslau and Marburg, at both of which his work justified the expectations that were realized at Heidelberg.

Some of the visitors to the unveiling ceremonies wandered to the cemetery of the old university town on the Neckar, and sought out the grave of the great scientist. The monument is plain and simple and bears a simple inscription. The mound is covered with rhododendrons. Not far away is the granite block, almost overgrown with ivy, which marks the last resting place of that other world renowned scientist, Victor Meyer, the worthy successor of the immortal Bunsen.

A SIMPLE METHOD FOR DETERMINING THE EQUIVALENT WEIGHT OF SODIUM.

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One of the laboratory exercises usually assigned students of general chemistry early in their course is the determination of the equivalent weights of some of the more common metals. In order that the results may be used subsequently in introducing

the idea of valence, zinc and aluminum are commonly chosen, their equivalent weights being obtained by collecting the hydrogen evolved when a given weight of the metal is dissolved in an acid.

In order that a more complete series may be obtained, it is evident that the equivalent weight of such a metal as sodium should be added to the list. The decomposition of water by a given weight of this metal is the method naturally suggested, but owing to the difficulty of weighing accurately small quantities of sodium and to more or less danger connected with its use, the experiment is not usually attempted. One way out of the difficulty is to follow the old suggestion of Davy, that is, to make an amalgam with mercury, weigh out a definite amount of the amalgam, bring it in contact with water in a suitable vessel, collect the hydrogen, and determine the weight of the sodium by drying and weighing the residual mercury. The weight of the sodium is found by taking the difference between the weight of the amalgam and the weight of the mercury recovered.

A much more convenient method is the one described below. It has the merit of simplicity, it is easily manipulated, is quite accurate, free from danger, and can be completed in a few minutes.

Select a smooth piece of ordinary tin foil about 4 cm. square and weigh it accurately. Cut off a piece of sodium approximately 5 mm. square and 1 mm. thick, press it quickly two or three times between filter paper, trim off any corroded edges with a sharp knife and quickly wrap it tightly in the tin-foil, pressing down the corners so as to exclude all air. Weigh the tin-foil with the sodium. The difference between this and the previous weight gives the weight of the sodium.

Fill a 100 cc. eudiometer with distilled water, invert it in a 100 cc. porcelain dish and clamp it so that its mouth is about 1 cm. above the bottom of the dish. Slip the tin-foil containing the sodium under the mouth of the eudiometer, withdraw most of the water from the dish, and with a sharp pointed instrument (file) slightly bent near the sharp end, make a small puncture in the tin-foil. In a minute or two make another puncture if necessary. After it is certain that all the sodium has reacted with the water, determine the volume of the hydrogen in the usual way, calculate its weight and from it calculate the equivalent weight of sodium. The result will be found quite satisfactory.