THE STANDARDIZING OF PERMANGANATE OF POTASH, BY MEANS OF PLUMBIC OXALATE.

BY PROFESSOR F. STOLBA.*

INSTEAD of using oxalic acid the author prefers its lead salt, for the following reasons :—

1. It does not contain water, and can always be got pure.
2. It does not take up any moisture even after several days' exposure to the air.
3. Its high molecular weight (one part of it = $\cdot 427227$ parts of oxalic acid), also gives it an advantage over oxalic acid, as good results may be obtained even by using a less delicate balance.

The process of standardizing is conducted in exactly the same manner as when oxalic acid is used. Dilute sulphuric acid is added, and the liquid heated; plumbic sulphate and free oxalic acid are formed. One gram of plumbic oxalate corresponds with

$\cdot 379757$ Iron.
 $\cdot 189878$ Calcic Oxide.

The author gives the results of four experiments which agree exceedingly well, different quantities of plumbic oxalate being taken each time for the standardizing. The following process is used by the author to prepare pure plumbic oxalate:—Plumbic acetate is dissolved in water, acidified with acetic acid, the solution boiled with lead foil to throw down traces of copper or silver, then precipitated with oxalic acid in excess. This precipitate is washed by decanting with water till the washings are no longer acid, then dried at 120°C . and kept in a well-closed bottle.

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